

15. Which of the following is not useful in handling acids?
- If corrosions acids are swallowed, neutralize it by NaOH
 - Wear goggles, gloves and laboratory coat
 - If an acid enters your eye, wash with water repeatedly
 - If concentrated acid is spilled on to your body, wash it with running water and 10% Na_2CO_3 .
16. Which of the following salt is given before gastro intestinal x – ray photography?
- NaCl
 - BaSO_4
 - $(\text{NH}_4)_2\text{CO}_3$
 - COCl_2
17. Which of the following is the property of base?
- They have sour taste
 - Their POH is greater than 7
 - Molten bases conduct electricity
 - They displace hydrogen from dilute acid
18. The pH of an orange juice was determined to be 5. What is the hydroxide ion concentration in the juice?
- $1 \times 10^{-5}\text{m}$
 - $1 \times 10^{-11}\text{m}$
 - $1 \times 10^{-2}\text{m}$
 - $1 \times 10^{-9}\text{m}$
19. $\text{NaOH} + \text{ZnO} \rightarrow \text{O}_2 + ? + ?$
- $\text{ZnNaO}_2 + \text{H}_2\text{O}_2$
 - $\text{ZnNaO}_2 + \text{H}_2$
 - $\text{Na}_2 \text{ZnNa}_2 + \text{H}_2\text{O}$
 - $\text{ZnNa} + \text{H}_2\text{O}$
20. Which of the following cannot be used to prepare base?
- The reaction between soluble base and soluble salt.
 - Direct combination of metal with hydrogen and oxygen.
 - By the reaction of highly reactive metal with water.
 - The reaction between basic an hydride and water.
21. A solution has a P^{H} of 9. The solution is best described as:
- Strongly alkaline
 - Weakly acidic
 - weakly alkaline
 - Strongly acidic
22. Which one of the following is NOT the property of bases?
- Having PH value greater than 7
 - Releasing hydroxide ions in water solution
 - Changing blue litmus paper to red
 - Having bitter test
23. Which one of the following is insoluble salt?
- AgCl
 - AgNO₃
 - NaCl
 - KNO₃
24. Which of the following mixtures has the highest PH?
- rain water
 - Vinegar
 - milk of magnesia
 - ambo water
25. What mass of NaOH is present in 10ml of 4M solution of sodium hydroxide?
- 16g
 - 0.16g
 - 1.6g
 - 160g
26. A solution A and B have PH values of 2 and 10 respectively. How many times is the H⁺ ion greater concentration in solution A than that of solution B?
- 3
 - 5
 - 4
 - 6
27. Which one of the following is a Lewis Base?
- H⁺
 - AlCl₃
 - CO₂
 - H₂O
28. Which of the following is a Lewis base?
- BF₃
 - NH₃
 - HF
 - AlCl₃
29. A ternary salt among the following is:
- $\text{Ca}(\text{HSO}_4)_2$
 - $\text{Ca}(\text{ClO}_4)_2$
 - Al_2S_3
 - $\text{CH}_3 \text{COONa}$
30. Which one of the following substance is used in dry cleaning?
- CCl₄
 - C₂Cl₄
 - C₆H₆
 - All of the above
31. Which one of the following bases is the strongest?
- Ba (OH)₂
 - Zn(OH)₂
 - Al(OH)₃
 - Mg(OH)₂

32. Which one of the following chemical processes does NOT form a hydroxide base?
- a) Reaction of an active metal with water
 - b) Reaction of an alkali with a solution of a salt
 - c) Direct combination of elements
 - d) Reaction of a soluble metallic oxide with water
33. Most of the properties of water solution of a base are associated to:
- a) The hydrogen ions in the solution
 - b) The charges of the ions in the solution
 - c) The metal ions in the solution
 - d) The hydroxide ions in the solution
34. Heating of a mixture of a salt and an alkali evolved a gas that forms a dense white fume with hydrochloric acid. The original salt could be:
- a) An ammonium salt
 - b) A nitrate
 - c) A sulphide
 - d) A carbonate
35. Which one of the following salts is soluble in water?
- a) CaCO_3
 - b) AgCl
 - c) BaSO_4
 - d) Na_2CO_3
36. Which of the following group of salts give a white precipitate when a solution of barium chloride is added to them?
- a) Nitrates
 - b) Chlorates
 - c) Permanganates
 - d) Sulphates